



UNIVERSITY OF SWAZILAND
Faculty of Health Sciences
Department of Environmental Health Science

MAIN EXAMINATION PAPER 2019

TITLE OF PAPER : ORGANIC CHEMISTRY AND BIOCHEMISTRY FOR NURSES

COURSE CODE : GNS 112

DURATION : 2 HOURS

MARKS : 100

INSTRUCTIONS :

- READ THE QUESTIONS & INSTRUCTIONS CAREFULLY
- ANSWER **ANY FOUR** QUESTIONS
- EACH QUESTION **CARRIES 25** MARKS.
- WRITE NEATLY & CLEARLY
- NO PAPER SHOULD BE BROUGHT INTO OR OUT OF THE EXAMINATION ROOM.
- BEGIN EACH QUESTION ON A SEPARATE SHEET OF PAPER.

DO NOT OPEN THIS QUESTION PAPER UNTIL PERMISSION IS GRANTED BY THE INVIGILATOR.

QUESTION ONE

- a. Draw structures corresponding to the following IUPAC names
- (i) 2,3 dichloro-4-methylhexane [5 Marks]
 - (ii) 3-iodo-2,2,4,4 tetramethylpentane [6 Marks]
 - (iii) *N,N* diethylamine [4 Marks]
- b. What are the 4 structural levels of proteins? [4 Marks]
- c. Name the three processes by which organisms utilize energy. [6 Marks]
- Total:** [25 Marks]

QUESTION TWO

- a. How do the following factors affect a SN2 and SN1 reactions
- (i) Substrate [4 Marks]
 - (ii) Nucleophile [4 Marks]
 - (iii) Leaving group [4 Marks]
 - (iv) Solvent [4 Marks]
- b. What is the special class of lipids that has a characteristic ring system? Give one example of this class. [3 Marks]
- c. How do phospholipids organize themselves in cell membranes? [3 Marks]
- d. What type of bond links nucleotides together in DNA and RNA? [3 Marks]
- Total:** [25 Marks]

QUESTION THREE

- a. Give examples of the following types of proteins;
- (i) Regulatory proteins [2 Marks]
 - (ii) Structural proteins [2 Marks]
 - (iii) Enzymes [2 Marks]
 - (iv) Defense proteins [2 Marks]
 - (v) Transport [2 Marks]
 - (vi) Storage [2 Marks]
- b. Draw the bond that links amino acids together in proteins. [4 Marks]

- c. Use carbon's electron configuration to explain what is meant by 'excited state' of reacting atoms. [4 Marks]
- d. The reaction of 2-butanol with a strong base yields a mixture of elimination products. Draw all possible elimination products for this reaction and identify the major product. [5 Marks]
- Total:** [25 Marks]

QUESTION FOUR

- a. Match each enzyme subclass with its function;

Enzyme subclass	Function
Reductases	Transfer of phosphate
Lipases	Isomerization of chiral center
Dehydrogenases	Hydrolysis of phosphate
Kinases	Introduction of a double bond
Dehydrases	oxidation
Oxidases	Loss of water
Epimerases	reduction
Nucleases	Hydrolysis of an ester

[16 Marks]

- b. What is basal metabolic rate? [3 Marks]
- c. What is the difference between anabolism and catabolism? [4 Marks]
- d. Digestion of fats yields _____ and _____. [2 Marks]

Total:

[25 Marks]

QUESTION FIVE

- a. Define, with examples, what is meant by isomerism. [4 Marks]
- b. Draw structural formulas for dichlorinated propane, $C_3H_6Cl_2$. [6 Marks]
- c. Draw structures of the compounds described below and give the IUPAC name for each structure
- (i) 2-hydroxy-propyl propanoate.
- (ii) 2,3-dichlorophenol.

(iii) 4-bromo-2,4-dihydroxy pentanoic acid

[3×5 Marks]

Total:

[25 Marks]

General data and fundamental constants

Quantity	Symbol	Value
Speed of light	c	$2.997\ 924\ 58 \times 10^8 \text{ m s}^{-1}$
Elementary charge	e	$1.602\ 177 \times 10^{-19} \text{ C}$
Faraday constant	$F = N_A e$	$9.6485 \times 10^4 \text{ C mol}^{-1}$
Boltzmann constant	k	$1.380\ 66 \times 10^{-23} \text{ J K}^{-1}$
Gas constant	$R = N_A k$	$8.314\ 51 \text{ J K}^{-1} \text{ mol}^{-1}$ $8.205\ 78 \times 10^{-2} \text{ dm}^3 \text{ atm K}^{-1} \text{ mol}^{-1}$ $6.2364 \times 10 \text{ L Torr K}^{-1} \text{ mol}^{-1}$
Planck constant	h $\hbar = h/2\pi$	$6.626\ 08 \times 10^{-34} \text{ J s}$ $1.054\ 57 \times 10^{-34} \text{ J s}$
Avogadro constant	N_A	$6.022\ 14 \times 10^{23} \text{ mol}^{-1}$
Atomic mass unit	u	$1.660\ 54 \times 10^{-27} \text{ Kg}$
Mass		
electron	m_e	$9.109\ 39 \times 10^{-31} \text{ Kg}$
proton	m_p	$1.672\ 62 \times 10^{-27} \text{ Kg}$
neutron	m_n	$1.674\ 93 \times 10^{-27} \text{ Kg}$
Vacuum permittivity	$\epsilon_0 = 1/c^2 \mu_0$ $4\pi\epsilon_0$	$8.854\ 19 \times 10^{-12} \text{ J}^{-1} \text{ C}^2 \text{ m}^{-1}$ $1.112\ 65 \times 10^{-10} \text{ J}^{-1} \text{ C}^2 \text{ m}^{-1}$
Vacuum permeability	μ_0	$4\pi \times 10^{-7} \text{ J s}^2 \text{ C}^{-2} \text{ m}^{-1}$ $4\pi \times 10^{-7} \text{ T}^2 \text{ J}^{-1} \text{ m}^3$
Magneton		
Bohr	$\mu_B = e\hbar/2m_e$	$9.274\ 02 \times 10^{-24} \text{ J T}^{-1}$
nuclear	$\mu_N = e\hbar/2m_p$	$5.050\ 79 \times 10^{-27} \text{ J T}^{-1}$
g value	g_e	2.002 32
Bohr radius	$a_0 = 4\pi\epsilon_0 \hbar^2 / m_e e^2$	$5.291\ 77 \times 10^{-11} \text{ m}$
Fine-structure constant	$\alpha = \mu_0 e^2 c / 2h$	$7.297\ 35 \times 10^{-3}$
Rydberg constant	$R_\infty = m_e e^4 / 8h^3 c \epsilon_0^2$	$1.097\ 37 \times 10^7 \text{ m}^{-1}$
Standard acceleration of free fall	g	$9.806\ 65 \text{ m s}^{-2}$
Gravitational constant	G	$6.672\ 59 \times 10^{-11} \text{ N m}^2 \text{ Kg}^{-2}$

Conversion factors

1 cal	=	4.184 joules (J)	1 erg	=	$1 \times 10^{-7} \text{ J}$
1 eV	=	$1.602\ 2 \times 10^{-19} \text{ J}$	1 eV/molecule	=	96 485 kJ mol ⁻¹

Prefixes	f	p	n	μ	m	c	d	k	M	G
	femto	pico	nano	micro	milli	centi	deci	kilo	mega	giga
	10^{-15}	10^{-12}	10^{-9}	10^{-6}	10^{-3}	10^{-2}	10^{-1}	10^3	10^6	10^9

GROUPS

PERIODS	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18	
	IA	IIA	IIIB	IVB	VB	VIB	VIIIB	VIII	IX	X	IB	IIIB	IIIA	IVA	VA	VIA	VIIA	VIIIA	
1	1.008 H																		4.003 He
2	6.941 Li 3	9.012 Be 4											10.811 B 5	12.011 C 6	14.007 N 7	15.999 O 8	18.998 F 9	20.180 Ne 10	
3	22.990 Na 11	24.305 Mg 12											26.982 Al 13	28.086 Si 14	30.974 P 15	32.06 S 16	35.453 Cl 17	39.948 Ar 18	
4	39.098 K 19	40.078 Ca 20	44.956 Sc 21	47.88 Ti 22	50.942 V 23	51.996 Cr 24	54.938 Mn 25	55.847 Fe 26	58.933 Co 27	58.69 Ni 28	63.546 Cu 29	65.39 Zn 30	69.723 Ga 31	72.61 Ge 32	74.922 As 33	78.96 Se 34	79.904 Br 35	83.80 Kr 36	
5	85.468 Rb 37	87.62 Sr 38	88.906 Y 39	91.224 Zr 40	92.906 Nb 41	95.94 Mo 42	98.907 Tc 43	101.07 Ru 44	102.91 Rh 45	106.42 Pd 46	107.87 Ag 47	112.41 Cd 48	114.82 In 49	118.71 Sn 50	121.75 Sb 51	127.60 Te 52	126.90 I 53	131.29 Xe 54	
6	132.91 Cs 55	137.33 Ba 56	138.91 *La 57	178.49 Hf 72	180.95 Ta 73	183.85 W 74	186.21 Re 75	190.2 Os 76	192.22 Ir 77	195.08 Pt 78	196.97 Au 79	200.59 Hg 80	204.38 Tl 81	207.2 Pb 82	208.98 Bi 83	(209) Po 84	(210) At 85	(222) Rn 86	
7	223 Fr 87	226.03 Ra 88	(227) **Ac 89	(261) Rf 104	(262) Ha 105	(263) Unh 106	(262) Uns 107	(265) Uno 108	(266) Une 109	(267) Uun 110									

TRANSITION ELEMENTS

140.12 Ce 58	140.91 Pr 59	144.24 Nd 60	(145) Pm 61	150.36 Sm 62	151.96 Eu 63	157.25 Gd 64	158.93 Tb 65	162.50 Dy 66	164.93 Ho 67	167.26 Er 68	168.93 Tm 69	173.04 Yb 70	174.97 Lu 71
232.04 Th 90	231.04 Pa 91	238.03 U 92	237.05 Np 93	(244) Pu 94	(243) Am 95	(247) Cm 96	(247) Bk 97	(251) Cf 98	(252) Es 99	(257) Fm 100	(258) Md 101	(259) No 102	(260) Lr 103

*Lanthanide Series

**Actinide Series

() indicates the mass number of the isotope with the longest half-life.