UNIVERSITY OF SWAZILAND
BACHELOR OF SCIENCE

MAIN EXAMINATION 2015

TITLE OF PAPER	:	INTRODUCTORY PHYSICAL CHEMISTRY
COURSE NUMBER	:	C202
TIME	:	3 HOURS
INSTRUCTIONS	:	THERE ARE SIX QUESTIONS
	:	ANSWER ANY FOUR QUESTIONS
	:	BEGIN THE ANSWER TO EACH QUESTION ON A SEPARATE SHEET OF PAPER
	:	DATA SHEETS ARE PROVIDED WITH THIS EXAMINATION PAPER

DO NOT OPEN THIS PAPER UNTIL THE INVIGILATOR INSTRUCTS YOU TO DO SO.

Question 1 [25 Marks]

a) Describe in detail the scientific basis of the van der Waals equation. [10]

b) A real gas equation of state for a gas is given by:

$$P = RT(V_m - \beta)^{-1} - (\alpha / T)V_m^{-2} \quad (1)$$

i) Derive an expression for $V_{m,c}$, Tc and P_c. [6]
ii) Find an expression for the Boyle's temperature, T_B. [4]
iii) Estimate the temperature at which oxygen behaves as an ideal gas, T_B given the constants: α=1.748 L^2atm mol^{-2}K and β= 0.0345 L mol^{-1}. [2]
iv) Estimate the radii of real gas molecules using equation (1) for real gases given a critical molar volume of 250 cm^3mol^{-1} [2]

c) Using the critical point expressions for $V_{m,c}$, Tc and P_c find an expression or value for compressibility at the critical point, Z_c [1]

Question 2 [25 MARKS]

a) Using examples and/or diagrams compare and contrast ***Any Two*** of the following terms
i) reversible and irreversible expansion [5]
ii) path and state functions [5]
iii) change in internal energy and change in enthalpy [5]

b) 4 moles of butane occupies 24 L at 310 K.
i) Derive an expression for work in a reversible isothermal expansion. [5]
ii) Calculate the work done when the gas expands isothermally against a constant external pressure of 100 torr. [5]
iii) Calculate the work that would be done if the same expansion in b(ii) occurred in a series of equilibrium steps . [5]

[assume butane behaves as an ideal gas]

Question 3 [25 Marks]

a) Write short notes on **Any Two** of the following concepts.
i) Statistical view of entropy [5]
ii) Clausius inequality [5]
iii) Second law of thermodynamics [5]

For each concept include the origin or a short derivation showing its origin, an example where applicable and the role or implication of each of the concepts in thermodynamics.

b) 1.00 mol of perfect gas at 27°C is expanded isothermally from an initial pressure of 3.00 atm to a final pressure of 1.00 atm . Calculate *q*, *w*, ΔS_{sys}, ΔS_{surr} and ΔS_{tot} if the expansion is done:
(1) reversibly, and [5]
(2) against a constant external pressure of 1.00 atm. [5]
(3) adiabatically against a constant pressure of 1.00 atm. [5]

Question 4 [25 Marks]

a) (i) Calculate the enthalpy of formation of $N_2O_5(g)$ from the following data: [9]

$$2NO(g)+O_2(g) \rightarrow 2NO_2(g) \qquad \Delta H^\theta = -114.1\ kJmol^{-1}$$

$$4NO_2(g)+O_2(g) \rightarrow 2N_2O_5(g) \qquad \Delta H^\theta = -110.2\ kJmol^{-1}$$

$$N_2(g)+O_2(g) \rightarrow 2NO(g) \qquad \Delta H^\theta = +180.5\ kJmol^{-1}$$

(ii) Using the enthalpy of formation of $N_2O_5(g)$ obtained from a(i) calculate the change in internal energy for the formation of $N_2O_5(g)$ [6]

b) (i) Derive Kirrchoff's equation: [4]

$$\Delta H_r(T_2) = \Delta H_r(T_1) + \Delta_r C_p \Delta T$$

(ii) Predict the standard enthalpy of reaction at 80°C for the reaction: [6]

$$2NO_2(g) \rightarrow N_2O_4(g)$$

Refer to table and the data below:

	C_p J $mol^{-1}K^{-1}$	ΔH_f(298 K) /kJ/mol
$N_2O_4(g)$	77.28	+9.2
$NO_2(g)$	37.20	+33.2

Question 5 [25 Marks]

a) Write short notes on the following:
(i) Gibbs Free Energy [4]
(ii) Helmholtz Function [4]

Use equations, examples and Application to clarify your discussion.

b) Derive the integrated Gibbs-Helmholtz equation [5]

$$\frac{\Delta G_2}{T_2} - \frac{\Delta G_1}{T_1} = \Delta H\left(\frac{1}{T_2} - \frac{1}{T_1}\right)$$

starting from the fundamental thermodynamic equation $dG = VdP - SdT$

c) Given the reaction:

$$N_2(g)+3H_2(g) \rightarrow 2NH_3(g)$$

Calculate the change in Gibbs free energy ΔG^θ

i) at 298K [5]

ii) at 500K [5]

iii) Comment on the significance of the values obtained in (i) and (ii). [2]

Question 6 [25 Marks]

a) Draw sketch of the phase diagrams of carbon dioxide and water and explain briefly the differences in slopes and curvature of the liquid-solid and the liquid-gas boundaries, respectively. [15]

b) i) Starting from the Gibbs function G=H-TS, .derive the Clausius-Clapeyron equation for evaporation in the form

$$ln\frac{p_2}{p_1^*} = -\frac{\overline{\Delta H}_{VAP}}{R}\left(\frac{1}{T_2} - \frac{1}{T_1^*}\right).$$ [5]

ii) The boiling point of methanol is at 64.6°C and 760 mm Hg. Predict the boiling point of benzene at 1.2 atm pressure. [5]

Heat capacities at 25°C

	$C_{v,m}$ $JK^{-1}\ mol^{-1}$	$C_{p,m}$ $JK^{-1}\ mol^{-1}$
He, Ne, Ar, Kr, Xe	12.47	20.78
H_2	20.50	28.81
O_2	21.01	29.33
N_2	20.83	29.14
CO_2	28.83	37.14
NH_3	27.17	35.48
CH_4	27.43	35.74
N_2O_4		77.28
NO_2		37.20

F.P Depression, B.P. Elevation

Solvent	F.P °C	K_f °C kg mol^{-1}	B.P (°C, 101kNm^{-2})	K_b °C kg mol^{-1}
Water	0	1.86	100.0	0.52
Benzene	5.51	5.10	80.1	2.60
Acetic Acid	16.6	3.90	118.1	3.10
Cyclohexane	6.5	20.2	81.4	2.79
Camphor	177.7	40.0	205	-
Nitrobenzene	5.7	6.9	210.9	5.24
Ethanol	-177		78.5	1.22
Chloroform	-64		61.3	3.63

Third Law entropies at 25°C, S_m^θ/J K^{-1} mol^{-1}

Solids		Liquids		Gases	
Ag	42.68	Hg	76.02	H_2	130.6
C(gr)	5.77	Br_2	152.3	N_2	192.1
C(d)	2.44			O_2	205.1
Cu	33.4			Cl_2	223.0
Zn	41.6	H_2O	70.0	CO	197.67
I_2	116.7			CO_2	213.7
S(Rh)	31.9	HNO_3	155.6	HCl	186.8
				H_2S	205.6
AgCl	96.2	C_2H_5OH	161.0	NH_3	192.5
AgBr	104.6	CH_3OH	126.7	CH_4	186.1
$CuSO_45H_2O$	305.4	C_6H_6	49.03	C_2H_6	229.4
$HgCl_2$	144	CH_3COOH	159.8	CH_3CHO	265.7
Sucrose	360.2	C_6H_{12}	298.2		

Useful Relations

$(RT)_{298\cdot15K}$=2.4789 kJ/mol
$(RT/F)_{298\cdot15K}$=0.025 693 V

T/K:	100.15	298.15	500.15	1000.15
T/Cm^{-1}:	69.61	207.22	347.62	695.13

1mmHg=133.222 N m^{-2}
hc/k=1.438 78x10^{-2} m K

1atm	1 cal	1 eV	1cm^{-1}
=1.01325x10^5 Nm^{-2}	=4.184 J	=1.602 189x10^{-19} J	=0.124x10^{-3} eV
=760torr		=96.485 kJ/mol	=1.9864x10^{-23}J
=1 bar		= 8065.5 cm^{-1}	

SI-units:

1 L = 1000 ml = 1000cm^3 = 1 dm^3
1 dm = 0.1 m
1 cal (thermochemical) = 4.184 J
dipole moment: 1 Debye = 3.335 64x10^{-30} C m
force: ***1N=1J m^{-1}= 1kgms^{-2}***=10^5 dyne pressure: ***1Pa=1Nm^{-2}***=1Jm^{-3}
1J = 1 Nm
power: 1W = 1J s^{-1} potential: 1V =1 J C^{-1}
magnetic flux: 1T=1Vsm^{-2}=1JCsm^{-2} current: 1A=1Cs^{-1}

Prefixes:

p	n	m	m	c	d	k	M	G
pico	nano	micro	milli	centi	deci	kilo	mega	giga
10^{-12}	10^{-9}	10^{-6}	10^{-3}	10^{-2}	10^{-1}	10^3	10^6	10^9

General Data

speed of light	**c**	**2.997 925x10^8 ms^{-1}**
charge of proton	*e*	1.602 19x10^{-19} C
Faraday constant	F=L*e*	9.648 46x10^4 C mol^{-1}
Boltzmann constant	**k**	**1.380 66x10^{-23} J K^{-1}**
Gas constant	**R=Lk**	**8.314 41 J K^{-1} mol^{-1}**
		8.205 75x10^{-2} dm^3 atm K^{-1} mol^{-1}
Planck constant	**h**	**6.626 18x10^{-34} Js**
	$\hbar = \frac{h}{2\pi}$	1.054 59x10^{-34} Js
Avogadro constant	**L or N_{av}**	**6.022 14x10^{23} mol^{-1}**
Atomis mass unit	u	1.660 54x10^{-27} kg
Electron mass	**m_e**	**9.109 39x10^{-31} kg**
Proton mass	m_p	1.672 62x10^{-27} kg
Neutron mass	m_n	1.674 93x10^{-27} kg
Vacuum permittivity	$\varepsilon_o = \mu_o^{-1}c^{-2}$	8.854 188x10^{-12} J^{-1} C^2 m^{-1}
Vacuum permeability	μ_o	4πx10^{-7} Js^{2}C^{-2} m^{-1}
Bohr magneton	$\mu_B = e\hbar/2m_e$	9.274 02x10^{-24} JT^{-1}
Nuclear magneton	$\mu_N = e\hbar/2m_p$	5.05079x10^{-27} JT^{-1}
Gravitational constant	G	6.67259x10^{-11} Nm^{2}kg^{-2}
Gravitational acceleration	**g**	**9.80665 ms^{-2}**
Bohr radius	a_o	5.291 77x10^{-11} m

THE PERIODIC TABLE OF ELEMENTS

NON-METALS — METALLOIDS — METALS

Group	1	2	3	4	5	6	7	8	9	10	11	12	13	14	15	16	17	18
	IA	IIA	IIIB	IVB	VB	VIB	VIIB	VIIIB			IB	IIB	IIIA	IVA	VA	VIA	VIIA	VIIIA
Period 1	1 **H** 1.008																	2 **He** 4.003
2	3 **Li** 6.94	4 **Be** 9.01											5 **B** 10.81	6 **C** 12.01	7 **N** 14.01	8 **O** 16.00	9 **F** 19.00	10 **Ne** 20.18
3	11 **Na** 22.99	12 **Mg** 24.31											13 **Al** 26.9	14 **Si** 28.09	15 **P** 30.97	16 **S** 32.06	17 **Cl** 35.45	18 **Ar** 39.95
4	19 **K** 39.10	20 **Ca** 40.08	21 **Sc** 44.96	22 **Ti** 47.90	23 **V** 50.94	24 **Cr** 52.01	25 **Mn** 54.9	26 **Fe** 55.85	27 **Co** 58.71	28 **Ni** 58.71	29 **Cu** 63.54	30 **Zn** 65.37	31 **Ga** 69.7	32 **Ge** 72.59	33 **As** 74.92	34 **Se** 78.96	35 **Br** 79.91	36 **Kr** 83.80
5	37 **Rb** 85.47	38 **Sr** 87.62	39 **Y** 88.91	40 **Zr** 91.22	41 **Nb** 91.22	42 **Mo** 95.94	43 **Tc** 98.9	44 **Ru** 101.1	45 **Rh** 102.9	46 **Pd** 106.4	47 **Ag** 107.9	48 **Cd** 112.4	49 **In** 114.8	50 **Sn** 118.7	51 **Sb** 121.8	52 **Te** 127.6	53 **I** 126.9	54 **Xe** 131.3
6	55 **Cs** 132.9	56 **Ba** 137.3	71 **Lu** 174.9	72 **Hf** 178.5	73 **Ta** 180.9	74 **W** 183.8	75 **Re** 186.2	76 **Os** 190.2	77 **Ir** 192.2	78 **Pt** 195.1	79 **Au** 196.9	80 **Hg** 200.6	81 **Tl** 204.4	82 **Pb** 207.2	83 **Bi** 208.9	84 **Po** 210	85 **At** 210	86 **Rn** 222
7	87 **Fr** 223	88 **Ra** 226.0	103 **Lr** 257	104 **Unq**	105 **Unp**	106 **Unh**	107 **Uns**	108 **Uno**	109 **Une**									

Lanthanides	57 **La** 138.9	58 **Ce** 140.1	59 **Pr** 140.9	60 **Nd** 144.2	61 **Pm** 146.9	62 **Sm** 150.9	63 **Eu** 151.3	64 **Gd** 157.3	65 **Tb** 158.9	66 **Dy** 162.5	67 **Ho** 164.9	68 **Er** 167.3	69 **Tm** 168.9	70 **Yb** 173.0
Actinides	89 **Ac** 227.0	90 **Th** 232.0	91 **Pa** 231.0	92 **U** 238.0	93 **Np** 237.1	94 **Pu** 239.1	95 **Am** 241.1	96 **Cm** 247.1	97 **Bk** 249.1	98 **Cf** 251.1	99 **Es** 254.1	100 **Fm** 257.1	101 **Md** 258.1	102 **No** 255

*Numbers **below** the symbol indicates the **atomic masses**; and the numbers above the symbol indicates the atomic numbers.*

SOURCE: International Union of Pure and Applied Chemistry, I mills, ed., Quantities, Units, and symbols in Physical Chemistry, Blackwell Scientific publications, Boston, 1988, pp 86-98.

Standard molar enthalpies of formation at 298.15 K

	M_r	ΔH_f^θ/KJ/mol		M_r	ΔH_f^θ/KJ/mol
$H_2O(g)$	18.015	-241.8	$O_3(g)$	47.998	+142.7
$H_2O(l)$	18.015	-285.8	$NO(g)$	30.006	+90.2
$H_2O_2(l)$	34.015	-187.8	$NO_2(g)$	46.006	+33.2
$NH_3(g)$	17.031	-46.1	$N_2O_4(g)$	92.012	+9.2
$N_2H_4(l)$	32.045	+50.6	$SO_2(g)$	64.063	-296.8
$N_3H(l)$	43.028	+264.1	$H_2S(g)$	34.080	-20.6
$N_3H(g)$	43.028	+294.1	$SF_6(g)$	146.054	-1209
$HNO_3(l)$	63.013	-174.1	$HF(g)$	20.006	-271.1
$NH_2OH(s)$	33.030	-114.2	$HCl(g)$	36.461	-92.3
$NH_4Cl(s)$	53.492	-314.4	$HCl(aq)$	36.461	-167.2
$HgCl_2(s)$	271.50	-224.3	$HBr(g)$	80.917	+36.4
$H_2SO_4(l)$)	98.078	-814.0	$HI(g)$	127.912	+26.5
$H_2SO_4(aq)$	98.078	-909.3	$CO_2(g)$	44.010	-393.5
$NaCl(s)$	58.443	-411.0	$CO(g)$	28.011	-110.5
$NaOH(s)$	39.997	-426.7	$AL_2O_3(\alpha,s)$	101.945	-1675.7
$KCl(s)$	74.555	-435.9	SiO2(s)	60.085	-910.9
$KBr(s)$	119.011	-392.2	$FeS(s)$	87.91	-100.0
$KI(s)$	166.006	-327.6	$FeS_2(s)$	119.975	-178.2
Diatomics(g)	—	**0**	$AgCl(s)$	143.323	-127.1

Enthalpies of fusion and evaporation ΔH_m/KJ/mol at the transition temperature

	T_f/K	Fusion[a]	T_b/K	Evaporation[b]	
He	3.5	0.021	4.22	0 084	
Ar	83.81	1.188	87.29	6.506	
H_2	13.96	0.117	20.38	0.9163	
N_2	63.15	0.719	77.35	5.586	
O_2	54.36	0.444	90.18	6.820	
Cl_2	172.12	6.406	239.05	20.410	
Br_2	265.90	10.573	332.35	29.45	
I_2	386.75	15.52	458.39	41.80	
Hg	234.29	2.292	629.73	59.296	
Ag	1234	11.30	2436	250.63	
Na	370.95	2.601	1156	98.01	
CO_2	217.0	8.33	194.84	25.23⊥	
H_2O	273.15	6.008	373.15	40.656	(44.016 at 298.15 K)
NH_3	195.40	5.652	239.73	23.351	
H_2S	187.61	2.377	212.80	18.673	
CH_4	90.68	0.941	111.66	8.18	
C_2H_6	89.85	2.86	184.55	14.7	
C_6H_6	278.65	10.59	353.25	30.8	
CH_3OH	175.25	3.159	337.22	35.27	(37.99 at 298.15K)

⊥ Sublimation: [a]various pressures: [b]at 1atm

Temperature dependence of heat capacities, $C_{p,m} = a+bT+cT^{-2}$

	a/J $K^{-1}mol^{-1}$	b/10^{-3}J $K^{-2}mol^{-1}$	c/10^5J $Kmol^{-1}$
Gases (298-2000K)			
He, Ne, Ar, Kr, Xe	20.78	0	0
H_2	27.28	3.26	0.50
O_2	29.96	4.18	-1.67
N_2	28.58	3.77	-0.50
Cl_2	37.03	0.67	-2.85
CO_2	44.23	8.79	-8.62
H_2O	30.54	10.29	0
NH_3	29.75	25.10	-1.55
CH_4	23.64	47.86	-1.92

Standard molar enthalpies of formation and combustion at 298.15 K.

	M_r	ΔH_f^θ/KJ/mol	ΔH_c^θ/KJ/mol
$CH_4(g)$	16.043	-74.81	
$C_2H_2(g)$	26.038	+226.8	1300
$C_2H_4(g)$	28.054	+52.30	1411
$C_2H_6(g)$	30.070	-84.64	1560
C_3H_6 cyclopropane(g)	42.081	53.35	2091
C_3H_6 (propene(g)	42.081	20.5	2058
C_4H_{10} n-butane (g)	58.124	-126.11	2877
C_5H_{12} n-pentane(g)	72.151	-146.4	3536
C_6H_{12} cyclohexane (l)	84.163	-156.2	3920
C_6H_{14} n-hexane (l)	86.178	-198.7	4163
C_6H_6 benzene (l)	78.115	+48.99	3268
C_8H_{18} n-octane (l)	114.233	-249.8	5471
$C_{10}H_8$ naphthalene (l)	128.175	+78.53	5157
CH_3OH (l)	32.042	-239.0	726.1
CH_3CHO (g)	44.054	-166.0	1193
CH_3CH_2OH (l)	46.070	-277.0	1368
CH_3COOH (l)	60.053	-484.2	874.5
$CH_3COOC_2H_5$ (l)	88.107	-486.6	2231
C_6H_5OH (s)	94.114	-165.0	3054
$C_6H_5NH_2$ (l)	93.129	-31.1	3393
$NH_2CO.NH$, urea(s)	60.056	-333.0	632.2
$CH_2(NH_2)CO_2H$, glycine (s)	75.068	-537.2	964.4
$C_6H_{12}O_6$, α-D-glucose (s)	180.159	-1274	2802
$C_6H_{22}O_6$, β-D-glucose (s)	180.159	-1268	2808
$C_{12}H_{22}O_{11}$, sucrose (s)	342.303	-2222	5645
$CH_3CH(OH)COOH$ lactic acid (s)	90.079	-694.0	1344

Source: American Institute of Physics handbook, McGraw-Hill

Standard molar Gibbs free energy and molar entropy of formation at 298.15 K

	M_r	ΔG_f^θ/KJ/mol	S^θ/J K^{-1} mol^{-1}		M_r	ΔG_f^θ/KJ/mol	S^θ/J K^{-1} mol^{-1}
$H_2O(g)$	18.015	-228.57	188.83	$O_3(g)$	47.998	163.2	238.93
$H_2O(l)$	18.015	-120.35	109.6	$NO(g)$	30.006	86.55	210.76
$H_2O_2(l)$	34.015	-120.35	109.6	$NO_2(g)$	46.006	51.31	240.06
$NH_3(g)$	17.031	-16.45	192.45	$N_2O_4(g)$	92.012	97.89	304.29
$N_2H_4(l)$	32.045	149.43	121.21	$SO_2(g)$	64.063	-300.19	248.22
$N_3H(l)$	43.028	327.3	140.6	$H_2S(g)$	34.080	-33.56	205.79
$N_3H(g)$	43.028	328.1	238.97	$SF_6(g)$	146.054	-1105.3	291.82
$HNO_3(l)$	63.013	-80.71	155.60	$HF(g)$	20.006	-273.2	173.78
$NH_2OH(s)$	33.030			$HCl(g)$	36.461	-95.30	186.91
$NH_4Cl(s)$	53.492	-202.87	94.6	$HCl(aq)$	36.461	-131.23	56.5
$HgCl_2(s)$	271.50	-178.6	146.0	$HBr(g)$	80.917	-53.45	198.70
$H_2SO_4(l)$)	98.078	-690.00	156.90	$HI(g)$	127.912	1.70	206.59
$H_2SO_4(aq)$	98.078	-744.53	20.1	$CO_2(g)$	44.010	-394.36	213.74
$NaCl(s)$	58.443	-384.14	72.13	$CO(g)$	28.011	-137.17	197.67
$NaOH(s)$	39.997	-379.49	64.46	Al_2O_3(□,s)	101.945	-1582.3	50.92
$KCl(s)$	74.555	-409.14	82.59	SiO_2	60.09	-856.64	41.84
$KBr(s)$	119.011	-380.66	95.90	$FeS(s)$	87.91	-100.4	60.29
$KI(s)$	166.006	-324.89	106.32	$FeS_2(s)$	119.975	-166.9	52.93
				$AgCl(s)$	143.323	-109.79	96.2
$He(g)$	4.003	0	126.15	$Hg(g)$	200.59	31.82	174.96
$Ar(g)$	39.95	0	154.84	$Hg(l)$	200.59	0	76.02
$H_2(g)$	2.016	0	130.684	$Ag(g)$	107.87	245.65	173.00
$N_2(g)$	28.013	0	191.61	$Ag(s)$	107.87	0	42.55
$O_2(g)$	31.999	0	205.138	$Na(g)$	370.95	76.76	153.71
$O_3(g)$	47.998	163.2	238.93	$Na(s)$	22.99	0	51.21
$Cl_2(g)$	70.91	0	223.07				
$Br_2(g)$	159.82	3.110	245.46				
$Br_2(l)$	159.82	0	152.23				
$I_2(g)$	253.81	19.33	260.69				
$I_2(s)$	253.81	0	116.135				

	M_r	ΔG_f^θ/KJ/mol	S^θ/J K^{-1} mol^{-1}
organic compounds			
$CH_4(g)$ methane	16.043	-50.72	186.26
$C_2H_2(g)$ ethyne	26.038	209.20	200.94
$C_2H_4(g)$ ethene	28.05	68.15	219.56
$C_2H_6(g)$ ethane	30.070	-32.82	229.60
C_3H_6 cyclopropane(g)	42.081	104.45	237.55
C_3H_6 propene(g)	42.081	62.78	267.05
C_4H_{10} n-butane (g)	58.124	-17.03	310.23
C_5H_{12} n-pentane(g)	72.151	-8.20	348.40
C_6H_{12} cyclohexane (l)	84.163	26.8	
C_6H_{14} n-hexane (l)	86.178		204.3
C_6H_6 benzene (l)	78.115	124.3	173.3
C_6H_6 benzene (g)	78.115	129.72	269.31
C_8H_{18} n-octane (l)	114.233	6.4	361.1
$C_{10}H_8$ naphthalene (l)	128.175		
CH_3OH (g)	32.042	-161.96	239.81
CH_3OH (l)	32.042	-166.27	126.8
CH_3CHO (g)	44.054	-128.86	250.3
CH_3CH_2OH (l)	46.07	-174.78	160.7
CH_3COOH (l)	60.053	-389.9	159.8
$CH_3COOC_2H_5$ (l)	88.107	-332.7	259.4
C_6H_5OH (s)	94.114	-50.9	146.0
$C_6H_5NH_2$ (l)	93.129		
$CH_2(NH_2)CO_2H$, glycine (s)	75.068	-373.4	103.5
$C_6H_{12}O_6$, □-D-glucose (s)	180.159		
$C_6H_{22}O_6$, □-D-glucose (s)	180.159	-910	212
$C_{12}H_{22}O_{11}$, sucrose (s)	342.303	-1543	360.2
$CH_3CH(OH)COOH$	90.079		
lactic acid (s)			

Source: American Institute of Physics handbook, McGraw-Hill.